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PROPERTIES *Pre-*

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part 26 (Colligative
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Elevation of Boiling*

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**12 XII What's the
Difference Between
Molarity and Molality?**

Chemistry ??? ?????

?? ????? ?????? ?? ?????

**??? ?? ????? ?? (how
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PROPERTIES-4 09

Solution Part 09

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Explained *LTHS*

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Concentration Dilution

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Breaking up the
solute into individual
components
(expanding the solute)

Step 2 Overcoming
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Step 1 Breaking up
the solute into
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...

Colligative Properties
of Solutions.

Colligative Properties
Colligative Property: A
property that depends
only upon the number
of solute particles
(concentration), and
NOT upon their
identity. Three

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Important Colligative

Properties of
Solutions. Vapor-
pressure lowering

Boiling-point elevation

Freezing-point

depression Vapor-

Pressure Lowering

Vapor pressure: is the

pressure exerted by a

vapor that is in

dynamic equilibrium

with its liquid

(molecules are

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moving back and forth

Properties Of

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Title: Colligative

Properties of

Solutions 1 Colligative

Properties of

Solutions. Boiling Pt

Elevation and

Freezing Pt

Depression; 2

Colligative Properties.

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depend only on the number of solute particles in a solution ; does not depend on the identity of particles; 3 Boiling Point Elevation. a nonvolatile solute lowers the vapor pressure

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Colligative Properties of Solutions. Jacobus Henricus van 't Hoff (1852-1911) Slide 2. Colligative Properties. Colligative properties are those that depend on the concentration of particles in a solution, not upon the identity of those particles. Boiling Point Elevation. Freezing

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Point Depression.
Osmotic Pressure.
Slide 3. Freezing
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Chemistry

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Colligative. NaCl
dissociates to form 2

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ion particles; It lowers
freezing points almost
twice as much as
methanol, a

nonelectrolyte

Colligative Properties.

It also includes a
worksheet in which
students can take
notes and practice the
skills learned during
the lesson..

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Author: wiki.ctsnet.org-

Kathrin Abendroth-20

20-10-08-17-06-20

Subject: Colligative
Properties Of Solution

Ppt Keywords: colligat
ive,properties,of,soluti

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1. Solutions

Colligative Properties

- Changes in colligative properties depend only on the number of solute particles present, not on the identity of the solute particles. •

Among colligative properties are Vapor

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pressure lowering
Boiling point elevation
Melting point
depression Osmotic
Pressure. 2.

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Solutions. Vapor-
pressure lowering
Boiling-point elevation
Freezing-point
depression Vapor-
Pressure Lowering
Vapor

Colligative Properties

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Both solutions have
the same freezing
point, boiling point,
vapor pressure, and

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osmotic pressure
because those
colligative properties
of a solution only
depend on the
number of dissolved
particles. Other non-
colligative properties
including e.g.
viscosity, surface
tension, and solubility
are different. 4.

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Effects of Solutes on
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Colligative Properties

- Chapter 19: Molality

and Colligative

Properties HW Ch. 19

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(on problems that are
a-z, please do a and b
only) Use the beakers
to demonstrate this |
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Properties
PowerPoint
presentation ...*

DILUTE AQUEOUS
SOLUTIONS e.g. 1 M
NaCl = 1 Mol NaCl/L

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= 31.449 g NaCl / 1 L
solution But: 1 L water
weighs 1.00 kg at 20
0C ? In dilute solution,
Molality ? Molarity

CONVERSIONS

BETWEEN

SOLUTION

PROPERTIES

RAOULT'S LAW In

Ideal Solutions: $P_1 =$

$X_1 P_1^0$ Note: $P_1^0 =$

Vapor Pressure of

Pure Solvent VAPOR

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PRESSURE OF
SOLVENT (P_1) vs.
MOLE FRACTION OF
SOLVENT (X_1)
ELEVATION OF
BOILING POINT BPT.

COLLIGATIVE PROPERTIES

Among colligative
properties are Vapor
pressure lowering
Boiling point elevation
Melting point

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depression Osmotic
pressure © 2009,
Prentice-Hall, Inc.
Vapor Pressure

Because of solute-
solvent intermolecular
attraction, higher
concentrations of
nonvolatile solutes
make it harder for
solvent to escape to
the vapor phase. ©
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$= X_A P_A^*$ where X_A is the mole fraction of compound A, and P_A^* is the normal vapor pressure of A at that temperature.

*Chapter 13 Properties
of Solutions - Colby
College*

Colligative Properties
Of Solutions. 1.

Colligative Properties
of Solutions. 2.

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The presence of solutes affects the properties of solutions. Some of these properties are not dependant on what dissolved substance but only on how much. Colligative properties are those that depend on the concentration of solute particles but

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not on their
identity.
Colligative
Properties?
Properties Ppt

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Of Solutions -
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Colligative Properties
of Ionic Solutions
Colligative properties
of solutions depends
upon the total
concentration of

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particles. Each equation describing colligative properties must be modified to account for this with ionic solutions since each ionic compound gives more than one mole of ions for every mole of compound.

PowerPoint

Presentation

Colligative Properties.

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The properties of the solutions which depend only on the number of solute particles but not on the nature of the solute are called Colligative properties. The four important colligative properties are: (i) Relative lowering in vapour pressure (ii) Elevation in boiling point (iii)

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Depression in
freezing point (iv)
Osmotic pressure.

*Colligative Properties
| Chemistry, Class 12,
Solutions*

Apr 24, 2020 - By
Nora Roberts # PDF
Solutions And
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Ppt # solution and
colligative properties
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are homogeneous mixtures of two or more pure substances one constituent is usually regarded as the solvent and the others as solutes in a solution the solute is

*Solutions And
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Colligative Properties
(solutions) A.

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Definition Colligative
Property property that
depends on the
concentration of
solute particles, not
their identity 4
Colligative
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