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of how the atoms and molecules are arranged in solids, liquids, and gases. A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces

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1) calculate moles of each element (if given mass) 2) divide each mole quantity by the smallest mole value from step 1. 3) use whole number values from step 2 as the subscripts to the original equation to write the empirical formula. Percent Composition. -the relative amounts of each element in a compound.

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